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Lewis structures for the following
molecules. Part c is done for you. a)
H₂O (oxygen is central) b) SO₂ c) N₂
It t-t (a [Jrt @ N=N @ @ f / @ 2 In
question 1 above you should see three
different kinds of bonds as

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Answers
best you can. a) single bond: one electron pair (e-p)
b) double bond:

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Structure of the Atom Note the following symbols: (they are not to scale) = proton (positive charge) = electron (negative charge) = neutron (no charge) The following three

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Answers are hydrogen atoms: 1 1H 2
1H 3 1H The following three
diagrams are carbon atoms: Notice the
type of notation used for atoms: 1 2
13 14 6C 6 C 6C (6 ...

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included in the Lewis Structure. $4 + 3(6) + 2 = 24$ (carbon has four and

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Answers
each oxygen has six; add two for the -2 charge) $5 + 3(1) = 8$ (nitrogen has five; each hydrogen has one) Step #2: Draw the “ skeleton structure ” with the central atoms and the other atoms, each connected with a single bond. Step #3: Add six more electron dots to each atom except the

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Steps for Writing Lewis Structures.

Example $\left(\frac{\text{PageIndex}\{2\}}{\wedge} \right) 1.$

Determine the total number of valence electrons in the molecule or ion. Each hydrogen atom (group 1) has one valence electron, carbon (group 14)

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Answers
has 4 valence electrons, and oxygen (group 16) has 6 valence electrons, for a total of $[(2)(1) + 4 + 6] = 12$ valence electrons. 2.

10.5: Writing Lewis Structures for Covalent Compounds ...

included in the Lewis Structure. 4 +

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$3(6) + 2 = 24$ (carbon has four and each oxygen has six; add two for the -2 charge) $5 + 3(1) = 8$ (nitrogen has five; each hydrogen has one) Step #2: Draw the “ skeleton structure ” with the central atoms and the other atoms, each connected with a single bond.

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