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~~Calculate the Molarity of~~

~~Acetic Acid in Vinegar~~

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Equations Step by Step

Practice Problems | How to

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and Molecular Compounds |

How to Pass Chemistry How to

Do Solution Stoichiometry

Using Molarity as a

Conversion Factor | How to

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(Redox) Reactions Step-by-

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~~Balancing Reactions~~

~~Predicting Products~~

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Laboratory Manual

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Reactions. 1. Double
Replacement Reactions.
Objectives. The objectives
of this lab are to: a)
Perform and observe the
results of a variety of
double replacement
reactions, b) Become
familiar with some of the
observable signs of these
reactions, c) Identify the
products formed in each of
these reactions, d) Write
balanced chemical equations
for each double replacement
reaction studied.

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The objectives of this lab are to: a) Perform and observe the results of a variety of double replacement reactions, b) Become familiar with some of the observable signs of these reactions, c) Identify the products formed in each of these reactions, d) Write balanced chemical equations for each double replacement reaction studied.

Double Replacement Reactions Laboratory Manual

Double Replacement Reactions
A double replacement
reaction has the form: $AB +$

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CD → AD + CB There are four different possible outcomes to a reaction such as this:

[1] Formation of a gas.

There are certain compounds which are unstable and decompose to water and a gas. Three common ones are H_2CO_3 , H_2SO_3 and NH_4OH .

They decompose like this: $\text{H}_2\text{CO}_3 \rightarrow \text{H}_2\text{O} + \text{CO}_2$

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Double Replacement Reactions
All double replacement
reactions have the general
form: $AB + CD \rightarrow AD + CB$
Reactions that can be
classified as double
replacements include
precipitation reactions,
neutralization reactions and
gas forming reactions.

10: Double Replacement Reactions (Experiment) - Chemistry ...

The general form of a double-
replacement (also called

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double-displacement)

reaction is: (11.9.1) $AB + CD \rightarrow AD + BC$ In this reaction, A and C are positively-charged cations, while B and D are negatively-charged anions. Double-replacement reactions generally occur between substances in aqueous solution.

11.9: Double Replacement Reactions - Chemistry LibreTexts

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Acid-Base - Also called neutralization reactions. These double replacement reactions occur when an acid and a base react to make a salt and water. $HA + BOH \rightarrow BA + H_2O$. $HNO_3(aq) + NaOH(aq) \rightarrow NaNO_3(aq) + H_2O(l)$
For today's experiment you will classify reactions as one of the main 5 types of chemical reactions.

Lab 6 Introduction | Chemistry I Laboratory Manual

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Use your observations above,
information in the
introduction, and reading to
classify the above chemical
reactions in the table
below. First: classify the
reaction as a main type:
either synthesis,
decomposition, combustion,
single replacement OR double
replacement. If no reaction

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occurred, say N/A. Record your choice in the space provided.

Lab 6 Worksheet | Chemistry I Laboratory Manual

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reactions between compounds
were done in order to
determine the equation and
description of a new
substance. During the lab,
each participant was given
drop bottles, spot plates.

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